1. Preface

“...what we accomplished...came through basic research without really knowing exactly how we were proceeding; we ultimately came to realize, step by step, that our basic research was leading to something really useful...and I think that’s what the Nobel Prize is all about: to do work that turns out to be useful to society in some way and certainly other fields in science.”

Professor Richard R. Schrock, my PhD advisor at MIT, said in the telephone interview right after the announcement of the 2005 Nobel Prize in Chemistry on 5 October 2005.

Basic research is valuable and can make significant impacts. Since the first reports published by my group in 2003, anilido PNP pincer complexes and their relevant derivatives have attracted increasing attention and gained popularity worldwide. This review, instead of having full coverage of what has been achieved to date, aims to address in its initial stage how these compounds were conceptually devised and gradually evolved along with my personal perspectives and tracks of this particular research.

2. Concept

Anilido or amido ligands are intrinsically $\sigma$- and $\pi$-donors that prefer electron-deficient metals to those carrying $\pi$-electrons for thermodynamic reasons. Phosphine ligands, on the other hand, are typically good $\sigma$-donors, with the possibilities of also being excellent $\pi$-acceptors if bearing strong electron-withdrawing substituents. Electron-rich metals, as a result, are bound to phosphine ligands more readily and strongly than anilido or amido ligands. In terms of the Hard and Soft Acids and Bases (HSAB) theory, hard N donors prefer high oxidation state early transition metals whereas soft P donors prefer low oxidation state late transition metals.

There have been a quite large number of amido (or anilido) complexes of early transition metals and phosphine complexes of late transition metals reported in the literature. Chemical bonds between mismatched HSAB pairs are thermodynamically unstable and kinetically labile. To exemplify, nickel amides or anilides are usually reactive bases or nucleophiles that deprotonate acids rapidly or react with electrophiles readily. The profound reactivity of these species is attributed to the inherent d$\pi$-p$\pi$ repulsion in the Ni-N bond. Compared with the ample examples of known zirconium amides or anilides,
phosphine complexes of zirconium are relatively rare. These mismatched donor-acceptor pairs are usually difficult to be stabilized thermodynamically or ligands in these pairs are sufficiently labile to kinetically dissociate readily, thereby easily resulting in coordinatively unsaturated species that transform subsequently or react rapidly.

Hybrid chelating ligands that contain both hard and soft donors are interesting as these chelates preserve the opportunities of creating simultaneously both matched and mismatched HSAB pairs upon coordination to a metallic element. While the matched pairs ensure thermodynamic stability of the derived metal complexes, the mismatched pairs endorse possible reactivity that is perhaps otherwise inaccessible. One remarkable example in this regard is the silyl bridged amido phosphine ligands (Figure 1) that have shown widespread reactivity with both hard and soft metals as pioneered by Fryzuk as of 1981. Of particular note in this particular system is the success of the derived metal complexes in reactivity with respect to dinitrogen activation and subsequent functionalization.

3. Evolution

Several drawbacks, however, are known in this silyl incorporated amido phosphine chemistry. With the embedded –SiMe₂CH₂– backbone that is inherently flexible, the soft phosphorus donors were found in some cases to dissociate readily from hard main group or transition metals. Undesirable degradation, on the other hand, may also occur to involve the amido phosphine ligand itself in certain complexes by means of the cleavage of the N-Si bonds or their adjacent C-H bonds, making the subsequent reactivity exploration somewhat unfeasible.

The employment of N-silyl substituents in amidometallic chemistry is popular. Side reactions involving the aforementioned bond cleavage, though undesirable, are not uncommon. Replacement of the silyl substituents in amides with other moieties in order to preclude these disadvantages has been documented. One successful method employing N-arylated instead of N-silylated triamidoamine complexes has been reported for catalytic reduction of dinitrogen to ammonia. Equally successful is another example concerning ortho-phenylene derived diamido/donor complexes for catalytic α-olefin polymerization, where a living fashion was found if the N-substituents are tert-butyl instead of trimethylsilyl groups. Though other options were also available, we were inspired by the triumph of the established triamidoamine as well as diamido/donor chemistry and attracted to anilido phosphine derivatives (Figure 1) because anilines are ubiquitous and inherently the ortho-phenylene backbone is thermodynamically much more robust and conformationally more rigid than the –SiMe₂CH₂– bridges in connecting N and P donors. These features are extremely crucial as the strong N-C(sp³) bonds are typically inert to enhance the possibilities of keeping the anilido phosphine ligands intact in the derived metal complexes during reactivity exploration and the rigid ortho-phenylene bridges sufficiently confine the N-C-C and C-C-P angles satisfactorily to diminish, to some extent, the propensities of donor arm dissociation.

4. Synthesis of anilido phosphine ligand precursors

As summarized in Scheme 1, the arylene bridged amido phosphine ligands are versatile, characteristic of having different hapticities, options of the third donor atom, varied bridges in connecting different donors, and distinct identities of substituents at these donors. In addition, alterations in anionic charges and incorporations of extra substituents into the arylene rings are also feasible.
Scheme 1. Representative examples of protio precursors of monoanionic anilido phosphine ligands
We reported in 2003 the syntheses of bis(2-diphenylphosphinophenyl)amine (H[1a])$^2$ and N-(2-diphenylphosphinophenyl)-2,6-diisopropylaniline (H[3b])$^3$ that were the first examples ever of precursors of any N-arylated anilido phosphine ligands. Relevant derivatives were soon developed to include monoanionic PNP 1 (Scheme 1a),$^{24-25}$ PNN 2 (Scheme 1b),$^{26}$ NP 3 (Scheme 1c),$^{27-28}$ and dianionic NPN,$^{29}$ etc. Coincidentally, Kaska et al. also independently published a bit later the preparation of H[1a] though the synthesis of its bis(2-bromophenyl)amine precursor by way of the Chapman rearrangement of an imidate as the key strategy requires much more experimental steps with quite harsher conditions thus characteristic of being comparatively more laborious, time consuming, and resulting in an extremely low overall yield.$^{30}$ Our approach outlined in Scheme 1 involves palladium catalyzed C-N bond-forming reactions followed by either nucleophilic or electrophilic phosphanylation, affording the protio ligand precursors H[1], H[2], and H[3] more efficiently in much higher yields.

Substituents at the two P donors in PNP 1 can be intentionally varied as exemplified by 1d.$^{25}$ Such desymmetrization is beneficial as the two phosphorus donors could be electronically and sterically distinguished so as to finely tune the electronic and steric structures of the derived metal complexes and deliberately induce vacant or active sites at the metal center for subsequent reaction chemistry. In view of this, PNN 2 was devised as an example of different approaches.$^{26}$ Conformationally, the electronic and steric properties of 2 lie somewhere in between tridentate 1 and bidentate 3. Tethered with a more flexible ethylene bridge, the amino donor in 2 is in principle hemilabile upon coordination to a metal. All in all, the anilido phosphine ligands PNP 1, PNN 2, NP 3 and their corresponding relevant analogues have integrated as a library entity, allowing us to examine exploratory chemistry with judicious pairing these ligands with appropriate metals.

Following the same concepts, phenolate phosphines with distinct hapticities, anionic charges, and having extra substituents at the arylene bridges are also hybrid chelates, providing equally interesting coordination chemistry with both hard and soft metals. Examples of these ligands and chemistry derived thereafter can be found somewhere else.$^{31-36}$

5. Representative anilido phosphine complexes

These anilido phosphine ligands, as planned, bind to both hard and soft transition and main group metals. Figure 2 illustrates some representative examples.

Figure 2. Representative examples of anilido phosphine complexes
Lithium amides are often versatile and convenient starting materials employed in metathetical reactions with metal (pseudo)halides. Lithium complexes of 1, 2, and 3 were successfully prepared from the reactions of nBuLi with H[1], H[2], and H[3], respectively, in either hydrocarbon (pentane, toluene) or ethereal (Et2O, THF, DME) solutions.2-3,24,28,37-38 Upon formation from ethereal solutions, these lithium complexes may be isolated as either ethereal free or ethereal adducts, the number of coordinated Et2O, THF, or DME in which ranges from 0 to 2 depending on the identity of anilido X-ray diffraction studies.40 These species, except \( [2\text{Li}]_2 \), are mononuclear as indicated by X-ray diffraction studies. As a result, the lithium atom in these complexes is 4- or 5-coordinate. X-ray studies also confirm the coordination of the soft P donors in these species to the hard Li atoms, even though these bonds are constructed with mismatched HSAB pairs. These Li-P bonds also retain in solutions as indicated by solution \(^{7}\text{Li}\) and \(^{31}\text{P}\) NMR studies.

Zirconium chemistry of anilido phosphine complexes was explored. While \( ([1\text{b}]\text{ZrCl}_2(\mu-\text{Cl}))_2 \) is a chloride bridged dimer, \( [1\text{b}]\text{ZrMe}_3 \) and \( [1\text{b}]\text{Zr(CH}_2\text{SiMe}_3)_3 \) are mononuclear as elucidated by X-ray studies.39 Consistent with the steric bulkiness of the trimethylsilylmethyl ligands, one of the P donors in \( [1\text{b}]\text{Zr(CH}_2\text{SiMe}_3)_3 \) is forced to dissociate from Zr, making this species 5-coordinate. In contrast, the trimethyl \( [1\text{b}]\text{ZrMe}_3 \) is 6-coordinate while the chloride \( ([1\text{b}]\text{ZrCl}_2(\mu-\text{Cl}))_2 \) is 7-coordinate. Collectively, the mismatched Zr-P bonds are retained in these species. Derived from the bidentate anilido phosphine ligands, complexes \([3\text{a}]\text{ZrCl}_2 \) and \( [3\text{b}]\text{ZrCl}(\text{THF}) \), on the other hand, are 6-coordinate, also featuring the mismatched Zr-P bonds as confirmed by X-ray diffraction studies.40

A number of divalent group 10 metal complexes of 1, 2, and 3 have been prepared and subject to reactivity studies on bond-breaking and bond-forming reactions. X-ray studies reveal that these complexes are all 4-coordinate, having an approximately square planar geometry with the mismatched M-N (M = Ni, Pd, Pt) bond constructed, thus conformationally characteristic of being amido PNP pincer complexes. An exogenous geometry with the mismatched M-N (M = Ni, Pd, Pt) 4-coordinate, having an approximately square planar X-ray studies reveal that these complexes are all mononuclear as indicated by X-ray diffraction studies.39 Consistent with the steric bulkiness of the trimethylsilylmethyl ligands, one of the P donors in \([1\text{b}]\text{Zr(CH}_2\text{SiMe}_3)_3 \) is forced to dissociate from Zr, making this species 5-coordinate. In contrast, the trimethyl \([1\text{b}]\text{ZrMe}_3 \) is 6-coordinate while the chloride \( ([1\text{b}]\text{ZrCl}_2(\mu-\text{Cl}))_2 \) is 7-coordinate. Collectively, the mismatched Zr-P bonds are retained in these species. Derived from the bidentate anilido phosphine ligands, complexes \([3\text{a}]\text{ZrCl}_2 \) and \( [3\text{b}]\text{ZrCl}(\text{THF}) \), on the other hand, are 6-coordinate, also featuring the mismatched Zr-P bonds as confirmed by X-ray diffraction studies.40

6. Arene C-H activation

Benzene, toluene, and xylenes are inert molecules and often used as solvents for organic syntheses. These arenes, however, are reactive enough to undergo C(sp\(^2\))H bond cleavage under extremely mild conditions upon reactions with divalent nickel and platinum complexes of 1.

As aforementioned, \([1\text{a}]\text{PtMe} \) remains intact upon heating in benzene at 150 °C for 3 days under aerobic conditions. In the presence of one equiv of B(C\(_6\text{F}_5\))\(_3\), however, the same solution at 25 °C was found to afford in 31 hours \([1\text{a}]\text{PtPh} \) (Scheme 2), a consequence resulting from benzene C-H activation.44 Alternatively, \([1\text{a}]\text{PtOTf} \) also reacts with benzene at 110 °C or higher temperatures in the presence of aliphatic amines such as NET\(_3\), MeNC\(_2\), or 1,4-diazabicyclo[2.2.2]octane (DABCO) to generate quantitatively \([1\text{a}]\text{PtPh} \).
Nickel complexes of 1 are also reactive in this bond activation chemistry (Scheme 3). In the presence of one equiv of B(C₆F₅)₃, the hydride complex [1b]NiH reacts with benzene at 25 °C to give [1b]NiPh successfully, though not as clean as what is found for the aforementioned platinum chemistry. The desired [1b]NiPh is produced as a minor product (ca. 20% as judged by ⁳¹P NMR spectroscopy), accompanied inevitably by the major [1b]Ni(C₆F₅) throughout the reaction. The concomitant formation of [1b]Ni(C₆F₅) and [1b]NiPh is suggestive of the competitive occurrence of C₆F₅ transfer from boron to nickel and intermolecular benzene C-H activation. Similar results are also found for [1c]NiH. Interestingly, replacing Lewis acidic B(C₆F₅)₃ with AlMe₃ allows for both [1b]NiH and [1c]NiH to react cleanly with benzene at 25 °C, affording quantitatively [1b]NiPh and [1c]NiPh, respectively. Reactions employing toluene and xylenes also proceed similarly, producing quantitatively [1b]NiAr and [1c]NiAr (Ar = tolyl, xylyl) with the least sterically hindered C(sp²)-H bond being mainly or exclusively activated.

Scheme 2. Benzene C-H activation by [1a]PtX (X = Me, OTf)

Scheme 3. Benzene C-H activation by [1b]NiH
7. Cross-coupling catalysis

Scheme 4 summarizes a number of cross-coupling reactions catalyzed by nickel or palladium complexes of these anilido phosphine chelates. In particular, divalent nickel complexes of 1, 2, and 3 are all active catalyst precursors for Kumada couplings of Grignard reagents with phenyl halides. Dependent on the identity of the anilido phosphine ligands, iodo, bromo, and even chloro electrophiles are compatible. In general, the activities of these catalysts increase following the order [1]NiCl < [2]NiCl < [3]NiCl. Of particular note are reactions employing chloro electrophiles and alkyl nucleophiles that contain β-hydrogen atoms. Catalysis having up to 100% conversion with 98% yield is achieved.

![Kumada couplings](image)

**Kumada couplings**

\[
\begin{align*}
\text{RMgCl} & \quad + \quad \text{X} \quad \xrightarrow{\text{Kumada}} \quad \text{Ph} \\
\text{R} & = \text{alkyl, aryl} \\
\text{X} & = \text{Cl, Br, I} \\
\text{L} & = 1, 2, 3 \\
\text{THF, 25 or 60 °C, 12 h} & \quad \rightarrow \quad \text{up to 100% conversion} \quad \text{98% yield}
\end{align*}
\]

![Heck couplings](image)

**Heck couplings**

\[
\begin{align*}
\text{X} & = \text{Cl, Br, I} \\
\text{Y} & = \text{4-NMe}_2, 4-\text{OMe}, \text{H}, 2-\text{F}, 2-\text{CHO}, 2-\text{CHO}, 2-\text{C(O)Me}, 2-\text{C(O)Me}, 4-\text{NO}_2 \\
\text{MeNO}_2, \text{NMP, 160 °C} & \quad \rightarrow \quad \text{air and water compatible} \\
\text{TON} & \quad \text{up to} \quad 4.5 \times 10^7 \text{Pd}^{-1} \\
\text{TOF} & \quad \text{up to} \quad 1.1 \times 10^6 \text{Pd}^{-1} \text{h}^{-1}
\end{align*}
\]

![Suzuki couplings](image)

**Suzuki couplings**

\[
\begin{align*}
\text{X} & = \text{Cl, Br, I} \\
\text{Y} & = \text{4-NMe}_2, 2-\text{OMe}, 4-\text{OMe}, 2,4,6-\text{Pr}_3, 2,4,6-\text{Me}_3, 2-\text{Me}, 3,5-\text{Me}_2, 2-\text{Me}, \text{H}, 2-\text{Br}, 2-\text{F}, 4-\text{F}, 4-\text{C(O)Me}, 4-\text{NO}_2, \\
\text{K}_2\text{PO}_4, \text{H}_2\text{O}, \text{dioxane, 110 °C} & \quad \rightarrow \quad \text{air and water compatible} \\
\text{TON} & \quad \text{up to} \quad 6.8 \times 10^7 \text{Pd}^{-1} \\
\text{TOF} & \quad \text{up to} \quad 3.1 \times 10^6 \text{Pd}^{-1} \text{h}^{-1}
\end{align*}
\]

![Sonogashira couplings](image)

**Sonogashira couplings**

\[
\begin{align*}
\text{X} & = \text{alkyl, benzyl} \\
\text{Y} & = \text{4-NMe}_2, 2-\text{OMe}, 4-\text{OMe}, 2-\text{Me}, 3,5-\text{Me}_2, 2-\text{Me}, 4-\text{Me}, \text{H}, 2-\text{Br}, 2-\text{F}, 4-\text{F}, 4-\text{C(O)Me}, 4-\text{NO}_2, \\
\text{NEt}_3, \text{1,4-dioxane} & \quad \rightarrow \quad \text{25 to 110 °C} \\
\text{TON} & \quad \text{up to} \quad 1.0 \times 10^7 \text{Pd}^{-1} \\
\text{TOF} & \quad \text{up to} \quad 3.1 \times 10^6 \text{Pd}^{-1} \text{h}^{-1}
\end{align*}
\]

The pincer complex [1a]PdCl is a versatile catalyst precursor, capable of mediating catalytic Heck olefination, Suzuki arylation, and Sonogashira alkynylation of (hetero)aryl halides. Compounds [{[3b]Pd(μ-Cl)}_2 and [3b]PdCl(PCy3), on the other hand, are also active in catalytic Suzuki-type reactions. The characteristic stability of these complexes in the presence of water under aerobic conditions at elevated temperatures makes these catalysts very user-friendly due to their easy manipulation and storage. A number of functional groups are compatible with these name reactions, including those characteristic of being sterically hindered, electronically activated, electronically neutral, and electronically deactivated. Extremely high turnover
numbers and turnover frequencies have been found for Heck- and Suzuki-type catalysis. Employment of a trace amount (e.g., ppm) of these anilido phosphine complexes is thus sufficient for these catalytic reactions to operate satisfactorily. Of particular interest in the Suzuki couplings is also the effective construction of sterically encumbered tri-ortho-substituted biaryl complexes. Though [1a]PdCl and [3b]PdCl exhibit comparable activities in the production of 2,2’,4,6-tetramethylbiphenyl (87% and 82% yield, respectively), the former outperforms the latter for the generation of 2,4,6-triisopropyl-2’-methylbiphenyl (72% and 34% yield, respectively) under similar conditions. Contrasting with [1a]PdCl, complex [1b]PdCl that is characteristically more electron-releasing and stericly demanding exhibits substantially unsatisfactory activities in both Suzuki- and Sonogashira-type catalysis under identical conditions, highlighting significantly the profound P-substituent effects of these anilido phosphine complexes on these coupling reactions.

8. Conclusions

The origin and evolution of arylene bridged amido phosphine ligands and their subsequent metal complexes are described. Having hybrid characteristics, these ligands bind to both hard and soft metals as exemplified in this review by lithium, zirconium, and group 10 metals, respectively. With the rigid and robust arylene bridges in these ligands, the derived metal complexes are remarkably thermally stable. Of note are those also stable in the presence of water under aerobic conditions. These results are unusual taking into account the inherent stability and reactivity of mismatched HSAB pairs. Though markedly stable, these complexes are reactive enough to promote inert chemical bond cleavage and to mediate catalytic C-C bond-forming reactions. The exploration of this basic research is truly fun and rewarding.

References

(1) Transcript of the telephone interview with Professor Richard R. Schrock can be found from the website of the Nobel Prizes at https://www.nobelprize.org/nobel_prizes/chemistry/laureates/2005/schrock-telephone.html
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Lan-Chang Liang is a Distinguished Professor of Chemistry at National Sun Yat-sen University and a joint Professor of Medicinal and Applied Chemistry at Kaohsiung Medical University, Taiwan. He undertook his PhD studies in 1995-1999 at Massachusetts Institute of Technology with Professor Richard R. Schrock on early transition metal chemistry of complexes containing amido chelates. After a postdoctoral stay with Professor T. Daniel P. Stack at Stanford University studying functional molecular models of copper oxygenases, he began his independent career in 2000 at NSYSU where his research program focuses on the development and application of new mismatched coordination compounds, particularly those competent in inert chemical bond activation and subsequent catalytic functionalization. His work has been recognized by Thieme Chemistry Journals Award (2006), one of top international inorganic chemists under 40 years of age (Inorganica Chimica Acta 2007), and Chemical Society of Japan with Distinguished Lectureship Award (2008).